



Re-Accredited 'B++' 2.86 CGPA by NAAC

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**

University Campus, Udhna-Magdalla Road, SURAT - 395 007, Gujarat, India.

**વીર નર્મદ દક્ષિણ ગુજરાત યુનિવર્સિટી**

યુનિવર્સિટી કેમ્પસ, ઉદ્દના-મગદલા રોડ, સુરત - ૩૯૫ ૦૦૭, ગુજરાત, ભારત.

Tel : +91 - 261 - 2227141 to 2227146, Toll Free : 1800 2333 011, Digital Helpline No.- 0261 2368888

E-mail : info@vnsgu.ac.in, Website : www.vnsgu.ac.in

**-: પરિપત્ર :-**

યુનિવર્સિટી સંલગ્ન વિજ્ઞાન વિદ્યાશાખા હેઠળની તમામ કોલેજોનાં આચાર્યશ્રીઓને જણાવવાનું કે, શૈક્ષણિક વર્ષ ૨૦૨૫-૨૬ થી અમલમાં આવનાર T.Y.B.Sc. Chemistry Sem.-5 & 6 Major, Minor અને SEC નો પેટાસમિતિ દ્વારા તૈયાર કરવામાં આવેલ અભ્યાસક્રમ સંદર્ભે રસાયણશાસ્ત્ર વિષયની અભ્યાસ સમિતિની તા.૨૮/૦૩/૨૦૨૫ ની સભાના ઠરાવ ક્રમાંક:૦૨ થી કરેલ ભલામણ સ્વીકારી વિજ્ઞાન વિદ્યાશાખાની તા.૩૦/૦૪/૨૦૨૫ની સભાનાં ઠરાવ ક્રમાંક:૧૭ થી કરેલ ભલામણ સ્વીકારી એકેડેમિક કાઉન્સિલની તા.૫/૫/૨૦૨૫ ની સભાનાં ઠરાવ ક્રમાંક: ૮૪ થી મંજૂર કરેલ છે. જેનો અમલ કરવા આથી જાણ કરવામાં આવે છે.

બિડાણ: ઉપર મુજબ

ક્રમાંક:ઓથો./પરિપત્ર/સિલેબસ/૧૧૯૯૨/૨૦૨૫

તા.૨૬-૦૫-૨૦૨૫

*Wifer*  
કુલસચિવ

પ્રતિ,

- ૧) યુનિવર્સિટી સંલગ્ન વિજ્ઞાન વિદ્યાશાખા હેઠળની તમામ કોલેજોનાં આચાર્યશ્રીઓ.  
.....આપશ્રીની કોલેજના સંબંધિત શિક્ષકોને જાણ કરી અમલ કરવા સારું.
- ૨) ડીનશ્રી, વિજ્ઞાન વિદ્યાશાખા.
- ૩) પરીક્ષા નિયામકશ્રી, પરીક્ષા વિભાગ, વીર નર્મદ દ. ગુ. યુનિવર્સિટી, સુરત.  
.....તરફ જાણ તેમજ અમલ સારું.

# VEER NARMAD SOUTH GUJARAT UNIVERSITY



**UNDER GRADUATE PROGRAMME  
IN  
CHEMISTRY  
UNDER FACULTY OF SCIENCE  
3 (YEARS DEGREE) AND 4 (YEARS HONOURS)**

**T. Y. B. Sc. Sem. - V and Sem. - VI**

**VEER NARMAD SOUTH GUJARAT UNIVERSITY, SURAT**  
**UNDER GRADUATE PROGRAMME IN CHEMISTRY MAJOR/MINOR/SEC**  
**3 (YEARS DEGREE) AND 4 (YEARS HONOURS)**

**B. Sc. – Sem. – V and VI CHEMISTRY**

|                           |   |
|---------------------------|---|
| <b>Programme Outcomes</b> | <p><b>PO-01: <u>Scientific Knowledge &amp; Conceptual Understanding</u>:</b> Develop a strong foundation in scientific principles, theories, and concepts across disciplines, fostering interdisciplinary learning, advance knowledge, and problem-solving abilities.</p> <p><b>PO-02: <u>Analytical &amp; Critical Thinking</u>:</b> Apply critical thinking and analytical reasoning to evaluate scientific data, hypotheses, and real-world problems, leading to evidence-based conclusions.</p> <p><b>PO-03: <u>Research &amp; Inquiry-based Learning</u>:</b> Develop investigative skills through experimentation, data analysis and scientific inquiry to contribute to research and innovation.</p> <p><b>PO-04: <u>Laboratory &amp; Technical Skills</u>:</b> Gain hands-on experience with laboratory techniques, instrumentation, and computational tools relevant to scientific research and industry applications.</p> <p><b>PO-05: <u>Digital &amp; Computational Literacy</u>:</b> Utilize digital tools, computational techniques, and emerging technologies such as AI, bioinformatics, and statistical modelling to enhance scientific learning and problem-solving.</p> <p><b>PO-06: <u>Environmental &amp; Societal Responsibility</u>:</b> Understand the role of science in addressing environmental, health and societal challenges, promoting sustainability and ethical responsibility.</p> <p><b>PO-07: <u>Effective Communication &amp; Collaboration</u>:</b> Develop proficiency in scientific communication, both written and oral, for effective dissemination of knowledge while collaborating in multidisciplinary teams.</p> <p><b>PO-08: <u>Innovation &amp; Entrepreneurship</u>:</b> Foster an entrepreneurial mind-set by applying scientific knowledge for innovation, technology development, and industry-oriented applications. Develop sustainable solutions to address real-world challenges in research and environmental management.</p> <p><b>PO-09: <u>Lifelong Learning &amp; Professional Growth</u>:</b> Cultivate curiosity and adaptability for continuous learning, equipping students for higher education, research, and professional careers.</p> <p><b>PO-10: <u>Ethical Leadership &amp; Value-based Education</u>:</b> Develop leadership qualities, ethical values, and a sense of responsibility in applying science for societal progress, aligning with Indian knowledge systems and global perspectives.</p> |
|---------------------------|---|

|                                    |   |             |             |             |             |             |             |             |             |
|------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Programme Specific Outcomes</b> | <p><b>PSO1:</b> Develop a strong conceptual understanding of organic, inorganic, physical, and analytical chemistry, integrating interdisciplinary scientific principles for advanced problem-solving and research.</p> <p><b>PSO2:</b> Demonstrate hands-on laboratory and computational skills, performing experiments independently and in teams, applying digital tools, and effectively communicating scientific findings.</p> <p><b>PSO3:</b> Apply key chemical concepts, analytical reasoning, and computational techniques to solve real-world challenges, emphasizing research, innovation, and technology-driven learning.</p> <p><b>PSO4:</b> Utilize chemical methodologies and problem-solving strategies to address unfamiliar scientific challenges, fostering critical thinking, adaptability, and lifelong learning.</p> <p><b>PSO5:</b> Develop industry-relevant skills, enhancing employability through innovation, entrepreneurship, sustainable applications, and professional ethics in chemistry and allied sciences.</p> <p><b>PSO6:</b> Engage in research-driven learning, applying experimental techniques, computational methods, and scientific inquiry to contribute to cutting-edge advancements in chemistry.</p> <p><b>PSO7:</b> Attain the level of proficiency required for higher education, research, and interdisciplinary scientific careers, integrating lifelong learning, ethical practices, and leadership.</p> <p><b>PSO8:</b> Design innovative and sustainable solutions using chemical principles to tackle local and global scientific, societal, and environmental challenges responsibly.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping of PO and PSO</b>       |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
| <b>PO1</b>                         | ✓   |             | ✓           |             |             |             |             | ✓           |             |
| <b>PO2</b>                         | ✓   |             | ✓           | ✓           | ✓           |             | ✓           |             | ✓           |
| <b>PO3</b>                         | ✓   | ✓           | ✓           | ✓           |             |             | ✓           | ✓           |             |
| <b>PO4</b>                         |   | ✓           |             | ✓           | ✓           | ✓           | ✓           |             | ✓           |
| <b>PO5</b>                         |   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
| <b>PO6</b>                         | ✓   | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |
| <b>PO7</b>                         | ✓   | ✓           |             | ✓           | ✓           |             |             | ✓           | ✓           |
| <b>PO8</b>                         |   |             | ✓           |             | ✓           |             |             |             | ✓           |
| <b>PO9</b>                         | ✓   | ✓           |             | ✓           | ✓           |             |             | ✓           | ✓           |
| <b>PO10</b>                        |   |             |             |             |             |             | ✓           |             | ✓           |

# VEER NARMAD SOUTH GUJARAT UNIVERSITY

**Programme Name:** Bachelor of Science Sem. - V

**About Programme:**

**Teaching and Evaluation Scheme:**

| <b>Structure</b> |             |                                   |                                   |                 |                         |    |               |    |        |    |                |    |                |    |       |    |
|------------------|-------------|-----------------------------------|-----------------------------------|-----------------|-------------------------|----|---------------|----|--------|----|----------------|----|----------------|----|-------|----|
| Course Category  | Course code | Course title                      | Marksheet title in English        | Level of Course | Teaching hours per week |    | Exam duration |    | Credit |    | Internal marks |    | External Marks |    | Total |    |
|                  |             |                                   |                                   |                 | TH                      | PR | TH            | PR | TH     | PR | TH             | PR | TH             | PR | TH    | PR |
| Minor            | CH-MJ-501   | Inorganic Chemistry               | Inorganic Chemistry               | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Major            | CH-MJ-502   | Organic Chemistry                 | Organic Chemistry                 | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Major            | CH-MJ-503   | Physical Chemistry                | Physical Chemistry                | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Minor            | CH-ME-501   | Chemical Analysis                 | Chemical Analysis                 | 200             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Minor            | CH-ME-502   | Applied Industrial Chemistry      | Applied Industrial Chemistry      | 200             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| SEC              | CH-SEC-501  | Spectral and Industrial Chemistry | Spectral and Industrial Chemistry | -               | 01                      | 02 | 0.5           | 02 | 01     | 01 | 12             | 12 | 13             | 13 | 25    | 25 |

# VEER NARMAD SOUTH GUJARAT UNIVERSITY

Programme Name: Bachelor of Science Sem. - VI

About Programme:

Teaching and Evaluation Scheme:

## Structure

| Course Category | Course code | Course title         | Marksheet title in English | Level of Course | Teaching hours per week |    | Exam duration |    | Credit |    | Internal marks |    | External Marks |    | Total |    |
|-----------------|-------------|----------------------|----------------------------|-----------------|-------------------------|----|---------------|----|--------|----|----------------|----|----------------|----|-------|----|
|                 |             |                      |                            |                 | TH                      | PR | TH            | PR | TH     | PR | TH             | PR | TH             | PR | TH    | PR |
| Major           | CH-MJ-601   | Inorganic Chemistry  | Inorganic Chemistry        | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Major           | CH-MJ-602   | Organic Chemistry    | Organic Chemistry          | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Major           | CH-MJ-603   | Analytical Chemistry | Analytical Chemistry       | 300             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |
| Minor           | CH-ME-601   | General Chemistry    | General Chemistry          | 200             | 02                      | 04 | 01            | 04 | 02     | 02 | 25             | 25 | 25             | 25 | 50    | 50 |

[Subject Code-2303000303012001]

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**

**T. Y. B. Sc. – SEM – V CHEMISTRY (MAJOR)**

**PAPER - CH-MJ-501**

**Inorganic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

**As per NEP 2020**

**To be implemented from the Academic year 2025-26**

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-501  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Inorganic Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to provide an in-depth understanding of the historical and modern aspects of inorganic chemistry, focusing on the contributions of Prof. P.C. Ray to Indian metallurgy and chemistry, as well as advanced concepts in metal carbonyls and transition metal complexes.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | The course will equip students with knowledge of ancient Indian metallurgical techniques, metal extraction processes, and their relevance to modern chemistry while also developing a strong foundation in the bonding, structure, and properties of metal carbonyls and transition metal complexes.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental chemical concepts, historical metallurgical processes, and bonding theories. They will retain principles like ligand field theory and Jahn-Teller effects, preparing for advanced studies and research.</p> <p><b>CO2-Understanding:</b> Students will comprehend Prof. P.C. Ray's contributions and their relevance to modern chemistry. They will understand transition metal bonding and the role of historical metallurgy in scientific advancements.</p> <p><b>CO3-Application:</b> Students will apply ancient metallurgical knowledge to modern material science and extractive metallurgy. They will use ligand field theory and computational techniques to predict metal complex properties.</p> <p><b>CO4-Analysis:</b> Students will analyse Indian metallurgical advancements, bonding theories, and their implications for chemical reactivity. They will evaluate experimental techniques in metal extraction and transition metal chemistry.</p> <p><b>CO5-Evaluation:</b> Students will assess India's metallurgical heritage, Prof. P.C. Ray's impact, and the reactivity of transition metal complexes. They will critically evaluate experimental data and methodologies in chemical synthesis.</p> <p><b>CO6-Creation:</b> Students will integrate ancient and modern techniques for sustainable metal extraction and material design. They will innovate in transition metal complex synthesis and develop environmentally responsible chemical processes.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>   | ✓           | ✓           | ✓           |             |             | ✓           | ✓           | ✓           |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |

# CH-MJ-501 (Inorganic Chemistry)

Program Name: T. Y. B. Sc. Sem. – V

Syllabus effective from June 2025

|                 |  |                            |
|-----------------|--|----------------------------|
| <b>Unit – 1</b> |  |                            |
|                 | <p><b>Prof. P. C. Ray: A Pioneer in Reviving Ancient Indian Metallurgy and Chemistry</b></p> <p><b>Contribution:</b> Prof. Prafulla Chandra Ray, the "Father of Indian Chemistry," highlighted India's advanced metallurgical practices and linked ancient mining and metalworking techniques showcasing India's scientific achievements and inspiring modern research.</p> <ul style="list-style-type: none"><li>• Wootz steel: The rise and fall of a great Indian technology.</li><li>• Mining and ore extraction.</li><li>• Metal and metal working technology.</li><li>• Gold extraction process.</li><li>• Zinc production.</li><li>• Copper mining and extraction process.</li><li>• Extraction of copper for Ayurvedic purposes.</li><li>• Copper alloys, Mercury, Lead and silver.</li><li>• Iron and steel in India.</li></ul> | <b>15 h</b>                |
| <b>Unit – 2</b> |  |                            |
|                 | <p><b>[A] Metal carbonyls</b></p> <ul style="list-style-type: none"><li>• Definition.</li><li>• Classification of metal carbonyl.</li><li>• Nature of Bonding in metal carbonyls.</li><li>• Structure and bonding in Ni(CO)<sub>4</sub>, Fe(CO)<sub>5</sub>, Cr(CO)<sub>6</sub>, Fe<sub>2</sub>(CO)<sub>9</sub>, Mn<sub>2</sub>(CO)<sub>10</sub>, Co<sub>2</sub>(CO)<sub>8</sub>.</li></ul> <p><b>[B] Bonding in transition Metal Complexes</b></p> <ul style="list-style-type: none"><li>• Jahn Teller Theorem.</li><li>• Distortion in octahedral complexes.</li><li>• Ligand Field Theory.</li><li>• Molecular energy level diagram and magnetic properties for [CoF<sub>6</sub>]<sup>3-</sup>, [Co(NH<sub>3</sub>)<sub>6</sub>]<sup>3+</sup>, [FeF<sub>6</sub>]<sup>3-</sup>, [Fe(CN)<sub>6</sub>]<sup>3-</sup>.</li></ul>           | <b>07 h</b><br><b>08 h</b> |

## Reference books:

- 1) A History of Hindu Chemistry from the Earliest Times to the Middle of the Sixteenth Century A.D. by Prof. P. C. Ray, (1902) published by The Bengal Chemical and Pharmaceutical Works.
- 2) History of Science and Technology in Ancient India: Volume II – Metal Technology by Debiprasad Chattopadhyaya (1986) published by Firma KLM, Calcutta.
- 3) The Rustless Wonder: A Study of the Iron Pillar at Delhi
- 4) Introduction to Indian knowledge system concepts and Applications by B. Mahadevan, Vinayak Rajat Bhatt, Nagendra Pavana R. N.
- 5) Advanced Inorganic Chemistry by Cotton and Wilkinson, John Wiley.
- 6) Inorganic Chemistry by J. D. Lee.
- 7) Theoretical Inorganic Chemistry by Day & Selbin, Affiliated East West Publ. Pvt. Ltd.

- 8) Uni. Chemistry by B. H. Mohan.
- 9) Structural Inorganic chemistry by A. F. Wells.
- 10) Chemical Bonding – an introduction by Rawal, Patel & Patel.
- 11) Basic Inorganic Chemistry by Cotton and Wilkinson.
- 12) A Text book of Inorganic Chemistry by P. L. Soni.
- 13) Introduction to Inorganic Chemistry by Durrant and Durrant.
- 14) Modern Coordination Chemistry by R. L. Burdett and R. G. Wilkinson.
- 15) Inorganic Chemistry-Principles of structure and reactivity by J. E. Huheey, E. A. Keiter.

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – V CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-501**

**Inorganic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-501   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Inorganic Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to develop students' practical skills in inorganic qualitative analysis by systematically identifying cations and anions using classical and modern analytical techniques.  |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to handle and analyse inorganic chemicals, apply qualitative tests for the detection of various cations and anions, and interpret results to develop a strong foundation in practical inorganic chemistry.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Recall and list the ions used in inorganic qualitative analysis, such as chloride, bromide, nitrate, and others, and understand their properties and behavior in qualitative tests.</p> <p><b>CO2-Understanding:</b> Understand and explain the fundamental principles behind the identification of inorganic ions, focusing on their chemical reactions, solubility rules, and behavior in qualitative analysis.</p> <p><b>CO3-Application:</b> Apply qualitative analysis techniques to identify unknown inorganic ions in laboratory experiments, ensuring precise application of principles and techniques.</p> <p><b>CO4-Analysis:</b> Analyse and interpret the results of qualitative tests, assessing the reliability and accuracy of the data obtained from various inorganic ion reactions.</p> <p><b>CO5-Evaluation:</b> Evaluate the various methods used in qualitative analysis, comparing their effectiveness in identifying different inorganic ions, and propose improvements or alternatives.</p> <p><b>CO6-Creation:</b> Design a comprehensive experimental procedure for the qualitative analysis of an unknown inorganic sample, integrating scientific inquiry, creativity, and sustainable practices.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO2</b>   | ✓           |             | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO3</b>   |             | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   | ✓           |             |             | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>   |             | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |

# CHP-MJ-501 (Inorganic Chemistry)

Third Year B. Sc. Semester - V

Syllabus effective from June 2025

Total Credit: 02 (04 hours)

## Inorganic Qualitative Analysis:

### List of inorganic chemicals used for inorganic qualitative analysis

|                  |   |
|------------------|---|
| <b>Chloride</b>  | : Cu <sup>+2</sup> , Cd <sup>+2</sup> , Fe <sup>+3</sup> , Mn <sup>+2</sup> , Co <sup>+2</sup> , Ni <sup>+2</sup> , Ca <sup>+2</sup> , Ba <sup>+2</sup> , Sr <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup>                                       |
| <b>Bromide</b>   | : Sr <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup>   |
| <b>Iodide</b>    | : K <sup>+1</sup>   |
| <b>Nitrite</b>   | : Na <sup>+1</sup> , K <sup>+1</sup>  |
| <b>Nitrate</b>   | : Co <sup>+2</sup> , Ni <sup>+2</sup> , Ba <sup>+2</sup> , Sr <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup>  |
| <b>Sulphite</b>  | : Na <sup>+1</sup>  |
| <b>Sulphide</b>  | : Zn <sup>+2</sup> , Sb <sup>+3</sup> ,   |
| <b>Sulphate</b>  | : Cu <sup>+2</sup> , Cd <sup>+2</sup> , Al <sup>+3</sup> , Fe <sup>+2</sup> , Zn <sup>+2</sup> , Mn <sup>+2</sup> , Co <sup>+2</sup> , Ni <sup>+2</sup> , Mg <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup>                                       |
| <b>Carbonate</b> | : Cu <sup>+2</sup> , Cd <sup>+2</sup> , Zn <sup>+2</sup> , Mn <sup>+2</sup> , Co <sup>+2</sup> , Ni <sup>+2</sup> , Ca <sup>+2</sup> , Ba <sup>+2</sup> , Sr <sup>+2</sup> , Mg <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup>                    |
| <b>Phosphate</b> | : Cu <sup>+2</sup> , Al <sup>+3</sup> , Fe <sup>+3</sup> , Zn <sup>+2</sup> , Mn <sup>+2</sup> , Co <sup>+2</sup> , Ni <sup>+2</sup> , Ca <sup>+2</sup> , Ba <sup>+2</sup> , Sr <sup>+2</sup> , Mg <sup>+2</sup> , Na <sup>+1</sup> , K <sup>+1</sup> , NH <sub>4</sub> <sup>+1</sup> |
| <b>Borate</b>    | : H <sub>3</sub> BO <sub>3</sub>  |

Inorganic qualitative analysis of a mixture containing six radicals. The mixture may be soluble in water or dilute HCl or Concentrated HCl including chromate and borate.

Note: Candidate should perform the analysis of at least eight mixtures.

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

## T. Y. B. Sc. – SEM – V CHEMISTRY (MAJOR)

## PAPER - CH-MJ-502

## Organic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-502  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Organic Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to provide students with a comprehensive understanding of reaction mechanisms, aromaticity, alkaloids, and synthetic drugs, emphasizing both theoretical concepts and their practical applications in organic chemistry.  |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn the fundamental mechanisms of esterification, hydrolysis, and elimination reactions, analyse aromaticity using Huckel's rule, explore the structure and synthesis of alkaloids, and understand the classification, synthesis, and applications of important synthetic drugs.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental reaction mechanisms such as esterification, hydrolysis, pyrolytic elimination, and their specific pathways (<math>B_{AC^2}</math>, <math>A_{AC^2}</math>, <math>A_{AC^1}</math>, <math>A_{AL^1}</math>, Cope, and Chugaev).</p> <p><b>CO2-Understanding:</b> Students will explain the concept of aromaticity, including Hückel's rule, and differentiate between aromatic, non-aromatic, and anti-aromatic compounds (benzenoids and non-benzenoids).</p> <p><b>CO3-Application:</b> Students will apply knowledge of alkaloid structures, their classification, and structural determination methods to interpret the chemistry of nicotine and papaverine.</p> <p><b>CO4-Analysis:</b> Students will analyse the structure-activity relationships (SAR) of synthetic drugs, their classification based on pharmacological action, and their synthesis.</p> <p><b>CO5-Evaluation:</b> Students will evaluate different organic reaction mechanisms, comparing their efficiencies and selecting the most suitable synthetic pathways for specific drug molecules.</p> <p><b>CO6-Creation:</b> Students will design and propose synthetic strategies for pharmaceutical compounds, incorporating sustainable, innovative, and ethical practices in drug development.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO6</b>   |             | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |

# CH-MJ-502 (Organic Chemistry)

Program Name: T. Y. B. Sc. Sem. – V  
Syllabus effective from June 2025

| Unit – 1 |  |             |
|----------|--|-------------|
|          | <b>(A) Reaction Mechanism:</b><br>(a) Different types of mechanism for Esterification and Hydrolysis: $B_{AC}^2$ , $A_{AC}^2$ , $A_{AC}^1$ , $A_{AL}^1$ , (b) Mechanism of formation and hydrolysis of amides. (c) Pyrolytic elimination: Cope and Chugaev reaction. | <b>10 h</b> |
|          | <b>(B) Aromaticity:</b><br>Introduction to Aromaticity, Huckel's Rule, Aromatic Character of Arenes, Definition & Examples of Aromatic, Non-Aromatic, Anti-Aromatic Compounds (Benzenoids and Non-Benzenoids)  | <b>05 h</b> |
| Unit – 2 |  |             |
|          | <b>(A) Alkaloids:</b><br>The occurrence, Classification, General methods to determine their structure, Analytical and Synthetic evidence to prove the structure of Nicotine and Papavarine.  | <b>08 h</b> |
|          | <b>(B) Synthetic Drugs:</b><br>Classification, based on pharmacological action, synthesis and uses of Amylnitrate, Chloroquine, Pyrimethamine, Sulphadiazine, Diazepam, Lidocaine, Chlorpropamide, Dapsone, Isoniazide, 5-Fluoro Uracil.                             | <b>07 h</b> |

## Reference Books:

- Natural Products: Chemistry and Applications by Sujata V. Bhat
- Organic Chemistry of Natural Products by O.P. Agarwal
- The Alkaloids: Chemistry and Biology by Geoffrey A. Cordell
- Advanced Organic Chemistry: Reactions, Mechanisms, and Structure by Jerry March
- A Guidebook to Mechanism in Organic Chemistry by Peter Sykes
- Foye's Principles of Medicinal Chemistry by David A. Williams, Thomas L. Lemke
- Essentials of Medicinal Chemistry by Korolkovas and Burckhalter
- Mechanism and Structure in organic chemistry-Goulde. S.
- Reaction mechanism in organic chemistry by Mukhargy & Singh
- Principles of reaction mechanism in organic chemistry by Dharmaraha & Chawla
- Organic reaction mechanism by Bansal Tata Mac. Hill
- Organic Chemistry (Vol I & II) 6 th E<sup>dn</sup>, I. L. Finar.
- Organic Chemistry by Hendrickson, Cram & Hammond
- Organic Chemistry by Brown R. F.
- Organic Chemistry by Solomon W. Graham
- Principles of Organic Synthesis- R. O. C. Norman
- Basic Principles of Organic chemistry, by R. Y. Caserio, W. A. Benjamin
- May's Chemistry of synthetic Drugs by Dyson.
- Chemistry of drugs, Ener and Caldwell
- Synthetic drugs by Tyagi and Yadav.

- Synthetic Organic Chemistry by O. P. Agarwal
- Organic Chemistry by Morrison and Boyd.
- Chemistry of organic Natural Product Vol. I & II by O. P. Agarwal.
- Chemistry of synthetic drugs by Trivedi
- Principles of Medicinal Chemistry Vol. I & II by S. S. Kadam, K. R. Mahadik, K. G. Bothara (Nirali Prakashan)
- Medicinal Chemistry by Asuthosh kar 4/e
- Organic reactions & their mechanism by P. S. Kalsi, New age international publishers

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – V CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-502**

**Organic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-502   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Organic Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to develop students' practical skills in organic chemistry by providing hands-on experience in quantitative estimation techniques and chromatographic separation methods.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to determine the concentration and purity of organic compounds using estimation techniques and apply chromatography to separate amino acids, enhancing their analytical and laboratory skills in organic chemistry.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the principles and procedures involved in organic estimation techniques such as the determination of ketones, formaldehyde, aspirin purity, and saponification value.</p> <p><b>CO2-Understanding:</b> Students will explain the theoretical concepts behind organic estimation methods and chromatography techniques, emphasizing accuracy and reproducibility in chemical analysis.</p> <p><b>CO3-Application:</b> Students will perform organic estimation experiments and apply chromatographic techniques to separate amino acid mixtures, demonstrating laboratory precision and problem-solving skills.</p> <p><b>CO4-Analysis:</b> Students will analyse experimental data from organic estimations and chromatography, interpreting results to assess the purity, concentration, and composition of organic compounds.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the efficiency of different estimation techniques and chromatographic methods, comparing their advantages, limitations, and industrial relevance.</p> <p><b>CO6-Creation:</b> Students will design and propose innovative experimental strategies for organic estimations and chromatographic separations, incorporating sustainable practices and modern analytical tools.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>   |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO6</b>   |             | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |

# **CHP-MJ-502 (Organic Chemistry)**

**Third Year B. Sc. Semester - V**

**Syllabus effective from June 2025**

## **ORGANIC ESTIMATION (Any Four)**

1. Determination of amount of Ketone (Acetone)
2. Determination of saponification value of an oil.
3. Determination of percentage purity of Aspirin.
4. Determination of amount of Formaldehyde in given solution
5. Determination of amount of Ethyl acetate in the given solution
6. Determination of amount of glucose in the given solution

**(Instead of Sample weighing, solution to be given)**

## **CHROMATOGRAPHY**

Chromatography separation of amino acid mixture by ascending paper chromatography.

**(Any two)**

1. Glycine + Methionine
2. Alanine + Methionine
3. Alanine + Valine

**VEER NARMAD SOUTH GUJARAT UNIVERSITY****T. Y. B. Sc. – SEM – V CHEMISTRY (MAJOR)****PAPER - CH-MJ-503****Physical Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-503   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Physical Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to provide students with a deep understanding of thermodynamics, electrochemistry, and their applications, highlighting the contributions of Indian and global scientists in advancing these fields.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will explore fundamental thermodynamic principles, chemical potential, fugacity, and the third law of thermodynamics while learning about electrochemical cells, electrode potentials, and the thermodynamics of electrochemical reactions. The course integrates theoretical concepts with real-world applications, emphasizing their significance in energy conversion, materials science, and industrial processes.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental thermodynamic and electrochemical concepts, including the third law, chemical potential, fugacity, electrode potential, and electrochemical cells.</p> <p><b>CO2-Understanding:</b> Students will explain the contributions of Indian scientists like Jnan Chandra Ghosh, A.K. Mukherjee, and K.S. Venkateswarlu in thermodynamics and electrochemistry, highlighting their impact on modern scientific advancements.</p> <p><b>CO3-Application:</b> Students will explain the contributions of Indian scientists like Jnan Chandra Ghosh, A.K. Mukherjee, and K.S. Venkateswarlu in thermodynamics and electrochemistry, highlighting their impact on modern scientific advancements.</p> <p><b>CO4-Analysis:</b> Students will analyse the thermodynamics of electrochemical cells, entropy changes, and energy relationships to predict reaction spontaneity and equilibrium conditions.</p> <p><b>CO5-Evaluation:</b> Students will evaluate different electrochemical cell classifications, electrode potentials, and thermodynamic parameters, comparing their efficiencies and industrial applications.</p> <p><b>CO6-Creation:</b> Students will design and propose innovative approaches to thermodynamic and electrochemical applications, such as battery technology and sustainable energy solutions.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>  |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>  |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO6</b>  |             | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |

# CH-MJ-503 (Physical Chemistry)

Program Name: T. Y. B. Sc. Sem. – V  
Syllabus effective from June 2025

| Unit – 1  |      |
|---|------|
| <p><b>1. Ramnathan, Subrahmanyam Chandrashekhar</b></p> <ul style="list-style-type: none"><li>• <b>Contribution in thermodynamics and energy:</b> Chandrashekhar's work on stellar structure and evolution of stars also has tried to thermodynamics, as it involves the application of thermodynamic principles to stellar dynamics, including energy generation, radiation pressure, and gravitational collapse.</li></ul> <p><b>2. L. N. (Ludwig) Boltzmann and the Indian influence</b></p> <ul style="list-style-type: none"><li>• <b>Contribution in third law of thermodynamics:</b> Though not Indian, Ludwig Boltzmann's work on statistical mechanics and the molecular kinetic theory had a significant influence on Indian physicists. The application of statistical mechanics, which explains macroscopic thermodynamic behaviour in terms of microscopic particle motion, became a key area of study for many Indian scientists, especially in quantum thermodynamics and low-temperature physics.</li></ul> <p><b>3. Jnan Chandra Ghosh</b></p> <ul style="list-style-type: none"><li>• <b>Contribution in change in entropy and free energy:</b> Studied colloid thermodynamics and phase transitions. Conducted research on entropy changes and free energy in chemical reactions. Worked on thermodynamic stability in solution.</li></ul> <p><b>Thermodynamics and its Indian Relevance</b></p> <p>Thermodynamics is integral to many fields of physics, chemistry, engineering, and even biological systems. The contribution of these Indian scientists helps illustrate how the laws of thermodynamics apply across different areas of science and how innovations in the theoretical frameworks of thermodynamics can lead to groundbreaking advances in technology, cosmology, and material science.</p> <p><b>Chemical Potential and Fugacity in Thermodynamics</b></p> <ul style="list-style-type: none"><li>• System – Open, Closed, and Isolated.</li><li>• Chemical potential (<math>\mu</math>) or Partial molal free energy</li><li>• Derivation of Gibb's – Duhem equation</li><li>• Derivation of chemical potential (<math>\mu</math>) equation for an ideal gas</li><li>• Lewis fugacity concept, fugacity function and its physical significance</li><li>• Concept of Activity and activity coefficient</li><li>• Name of methods to determine fugacity</li><li>• Determination of fugacity by graphical method.</li><li>• Concept of activity and activity coefficient</li><li>• Standard state of solid, liquid and gas</li><li>• Numerical problems</li></ul> | 15 h |

|                                  |  |             |
|----------------------------------|--|-------------|
|                                  | <p><b>The Third Law of Thermodynamics</b></p> <ul style="list-style-type: none"> <li>• The Nernst Heat Theorem (NHT) and its limitation</li> <li>• Statement of third law by Plank, Lewis – Randall</li> <li>• Consequences of third law of thermodynamics</li> <li>• Relation between Residual entropy and Boltzmann constant</li> <li>• Determination of absolute entropy of gas, liquid and solid</li> <li>• Application, importance and exception of third law of thermodynamics</li> <li>• Numerical problems.</li> </ul>   |             |
| <b>Unit – 2 Electrochemistry</b> |  | <b>15 h</b> |
|                                  | <p><b>A. K. Mukherjee</b></p> <ul style="list-style-type: none"> <li>• <b>Contribution in electrochemical cell:</b> He contributed significantly to the understanding of electrochemical cells, electrochemical equilibrium, and related processes. His work laid the foundation for many studies in electrochemical energy conversion and storage system in India.</li> </ul> <p><b>K. S. Venkateswarlu</b></p> <ul style="list-style-type: none"> <li>• <b>Contribution in applied electro chemistry:</b> He conducted research on electrode potential, and ionic mobility. His work helped to improve understanding of ion transport in solutions. He contributed to applied electrochemistry in battery and fuel cell development.</li> </ul> <p><b>Electrodicts:</b></p> <ul style="list-style-type: none"> <li>• Concept of oxidation and reduction, electrochemical series</li> <li>• Definition of electrode, standard electrode, half-cell, single electrode potential</li> <li>• Sign convention of electrode potential</li> <li>• Galvanic cell with example of Daniel cell</li> <li>• emf of a cell and its measurements by voltmeter and potentiometer</li> <li>• Standard cell (Western cell)</li> <li>• Primary and secondary electrode, Standard hydrogen electrode [SHE] and Calomel electrode.</li> </ul> <p><b>Thermodynamics of electrochemical cell:</b></p> <ul style="list-style-type: none"> <li>• Flow diagram of electrochemical cell</li> <li>• Reaction in reversible cell and concept of thermodynamics reversibility</li> <li>• Reversible and irreversible electrochemical cell</li> <li>• Relation between free energy change and electrical energy</li> <li>• Prediction of spontaneity of cell reaction</li> <li>• Determination of standard free energy change and equilibrium constant</li> <li>• Temperature coefficient of emf of a cell</li> <li>• Entropy change and enthalpy change of cell reaction</li> </ul> <p><b>Classification of electrochemical cell:</b></p> <ul style="list-style-type: none"> <li>• Chemical and concentration cell</li> <li>• Liquid Junction Potential (LJP), Methods for elimination of LJP (salt bridge)</li> <li>• Derivation of emf of concentration cell with and without transference.</li> <li>• <b>Application of electrochemistry in battery:</b> Li ion cell and Ni-Cd cell.</li> <li>• Numerical problems.</li> </ul> |             |

## Reference Books:

1. Thermodynamics and chemical equilibria by Jnan Chandra Ghosh.
2. Physical chemistry by B.K. Sharma
3. Elemental physical chemistry by Glasston & Lewis
4. Advance physical chemistry by D.N. Bajpai
5. Essential of physical chemistry by Bahl, Tuli & Bahl
6. Physical chemistry by Gurdeep Raj
7. Physical chemistry by K.L. Kapoor
8. Physical chemistry by Puri, Sharma, and Pathania
9. Principles of stellar dynamics, New York: Dover A textbook of physical chemistry by A. S. Negi & Anand
10. Atmospheric Brown clouds by Veerbhadharan Ramanathan
11. Physical chemistry by P.L. Soni & O.P. Dharmraj
12. Physical chemistry by K.K. Sharma & L.K. Sharma
13. Industrial chemistry by B.K. Sharma
14. Electro chemistry by B.K. Sharma
15. Electrochemistry: Principles and applications by K. S. Venkateshwarlu
16. Modern Electrochemistry by J'om Bockris and Reddy
17. Physical chemistry by G.M. Barrow
18. Numerical problems by D.V.S Jain

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – V CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-503**

**Physical Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-503  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Physical Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to equip students with practical skills in physical chemistry by conducting experiments related to reaction kinetics, electrochemical analysis, and thermodynamic measurements using instrumental techniques.  |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will investigate reaction rates, measure optical activity using polarimetry, determine pH and buffer capacity, and analyse solutions using conductometry and potentiometry. These experiments will enhance their understanding of chemical kinetics, acid-base equilibria, and electrochemical principles through hands-on experience.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental concepts of reaction kinetics, electrochemistry, conductance, potentiometry, pH metry, and polarimetry.</p> <p><b>CO2-Understanding:</b> Students will explain the principles of reaction rate studies, conductometry, potentiometry, and polarimetry, along with their applications in chemical analysis.</p> <p><b>CO3-Application:</b> Students will apply experimental techniques to determine reaction rates, dissociation constants, solubility products, buffer capacities, and optical activity of substances.</p> <p><b>CO4-Analysis:</b> Students will analyse the kinetic parameters of chemical reactions and evaluate the properties of electrolytes using pH metry, conductometry, and potentiometry.</p> <p><b>CO5-Evaluation:</b> Students will evaluate experimental data for reaction kinetics, conductivity, and potentiometric titrations to determine unknown concentrations and dissociation constants.</p> <p><b>CO6-Creation:</b> Students will design and propose innovative experimental approaches to optimize reaction conditions, develop precise analytical techniques, and apply electrochemical methods to real-world problems.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>  |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>  |             | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
| <b>CO6</b>                          |   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |

# CHP-MJ-503 (Physical Chemistry)

Third Year B. Sc. Semester - V

Syllabus effective from June 2025

Total Credit: 02 (04 hours)

1. To investigate the rate of reaction between  $K_2S_2O_8$  and KI,  $[a=b]$
2. To investigate the rate of reaction between  $K_2S_2O_8$  and KI,  $[a\neq b]$
3. To investigate the rate of reaction between  $H_2O_2$  and KI.  $[a=b]$
4. **Polarimetry:** Determination of angle of rotation of given substance using three different dilutions and determination of concentration of unknown solution. Sugar, Glucose, Tartaric acid.
5. **pH metry:** To measure pH of different buffer solution and to study the buffer capacity.
6. **pH metry:** To determine the dissociation constant of weak acid  $[CH_3COOH]$  and weak base  $[NH_4OH]$  by different dilution.
7. **Conductometry:** To determine the amount of  $BaCl_2$  in the given solution using  $K_2CrO_4$  solution.
8. **Conductometry:** To determine the amount of  $NaCl$  in the given solution using  $AgNO_3$  solution.
9. **Conductometry:** To determine the amount of  $CH_3COOH$  in the given solution using  $NH_4OH$  solution.
10. **Potentiometry:** To determine the normality of  $HCl$  solution using 0.5 N  $NaOH$  solution.
11. **Potentiometry:** To determine the solubility and solubility product of sparingly soluble salt  $AgCl$  by the titration of  $AgNO_3$  and  $NaCl$ .

**Note: Any seven practicals including TWO chemical kinetics should be performed by the students.**

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

## T. Y. B. Sc. – SEM – V CHEMISTRY (MINOR)

## PAPER - CH-ME-501

## Chemical Analysis (Minor) (2 Credits Theory + 2 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-ME-501  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Chemical Analysis (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to introduce students to the fundamental principles of analytical chemistry, focusing on quantitative analysis, error minimization, and instrumental techniques such as conductometric titration.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn various analytical methods, data treatment techniques, and statistical approaches to ensure accuracy and precision in chemical analysis. Additionally, they will gain hands-on experience in conductometric titrations, understanding their principles, applications, and advantages over traditional indicator-based methods.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental concepts of analytical chemistry, including classical and instrumental methods, types of errors, and principles of conductometric titrations.</p> <p><b>CO2-Understanding:</b> Students will explain the advantages, limitations, and applications of different analytical techniques, statistical data treatment, and error minimization in chemical analysis.</p> <p><b>CO3-Application:</b> Students will apply quantitative and computational techniques to analyse experimental data, perform conductometric titrations, and determine accuracy and precision in results.</p> <p><b>CO4-Analysis:</b> Students will critically analyse experimental outcomes, assess statistical significance, interpret conductivity data, and optimize analytical methodologies for enhanced accuracy.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the reliability of analytical methods, compare conductometric titrations with classical techniques, and assess systematic errors' impact on chemical data interpretation.</p> <p><b>CO6-Creation:</b> Students will design and develop innovative analytical procedures, integrate digital tools for chemical data processing, and propose new applications of conductometric titrations.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO2</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO4</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO5</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
| <b>CO6</b>                          | ✓  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |

# Third Year B. Sc. Sem. – V

## CH-ME-501 (Chemical Analysis)

Syllabus effective from June 2025

| Unit – 1 |  |             |
|----------|--|-------------|
|          | <p><b>Introduction to analytical chemistry</b><br/>Chemical and Instrumental Analysis (advantages and disadvantages)<br/>Overview of methods used in Quantitative analysis (classification of classical and instrumental analysis), Factors affecting the choice of analytical methods (in brief), Step in quantitative analysis (Flow diagram), Analytical methods on the basis of Sample size (in brief), Sampling methods. Sampling in different physical states.</p> <p><b>Treatment of analytical data</b><br/><b>Significant figures and rules of computation.</b><br/>Error Definition, Types of errors: Determinates errors, indeterminate errors, constant and proportional errors. Define and explain the following terms – Accuracy and Precision, mean, median, deviation, average deviation, standard deviation, variance, coefficient of variation, relative mean deviation, range, absolute errors, relative errors. Minimization of determinates errors, Normal error curve. Rejection of result from a set of results, 2.5 d rule, 4.0 d rule and Q-test. (Problems based on above topics).</p> | <b>15 h</b> |
| Unit – 2 |  |             |
|          | <p><b>Conductometric titration</b><br/>Principle and types of conductometric titrations</p> <ol style="list-style-type: none"><li>Strong acid vs strong base</li><li>Strong acid vs weak base</li><li>Weak acid vs strong base</li><li>Weak acid vs weak base</li><li>Strong acid + weak acid vs strong base</li><li>Strong acid + weak acid vs weak base</li><li>Precipitation titration (i) <math>\text{BaCl}_2</math> vs <math>\text{K}_2\text{CrO}_4</math> (ii) <math>\text{AgNO}_3</math> vs <math>\text{NaCl}</math></li></ol> <p>Advantage of conductometric titration over indicator method.</p>  | <b>15 h</b> |

### Reference Books:

- Quantitative Analysis by R. A. Day & A. L. Underwood, 6<sup>th</sup> ed. Pub. Prentice Hall of India ltd.
- Vogel's Text Book Inorganic Quantitative Analysis, 6 th ed.
- Analytical Chemistry (Principles & Technique) by Lary G. Hargis.
- Fundamental of Analytical Chemistry by Skoog D. A. & West D. M.
- Holler F. J. Instrumental Methods of Analysis by B. K. Sharma
- Instrumental analysis by R. D. Braun Mc Graw Hill.
- Analytical Chemistry by Gary Christian Instrumental methods of chemical analysis
- Dr. H. Kaur. Pragati prakashan Meerut.
- College Analytical Chemistry by Mangaonkar, Teckchandani, Sathe, Ghalsasi, Jain (Himalaya Publication House)

**VEER NARMAD SOUTH GUJARAT UNIVERSITY****T. Y. B. Sc. – SEM – V CHEMISTRY Practical (MINOR)****PAPER - CHP-ME-501****Chemical Analysis (Minor) (2 Credits Theory + 2 Credits Practical)****As per NEP 2020****To be implemented from the Academic year 2025-26**

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-ME-501  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Chemical Analysis (Practical)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to develop students' practical skills in analytical chemistry by applying conductometric titration techniques for precise determination of acid and salt concentrations.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to perform conductometric titrations for strong and weak acids, acid mixtures, and precipitation reactions, enhancing their understanding of conductivity changes during titration and the advantages of instrumental analysis over traditional methods.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the fundamental principles of conductometric titrations, normality calculations, and conductivity changes during acid-base and precipitation reactions.</p> <p><b>CO2-Understanding:</b> Students will explain the theoretical concepts behind conductometric titrations, including ion conductivity, endpoint determination, and differences between strong and weak acids.</p> <p><b>CO3-Application:</b> Students will apply conductometric methods to determine the normality of strong and weak acids, analyse acid mixtures, and perform precipitation titrations.</p> <p><b>CO4-Analysis:</b> Students will critically interpret conductometric titration curves, assess conductivity variations, and compare experimental results with theoretical expectations.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the accuracy, precision, and reliability of conductometric titrations, comparing them with conventional volumetric titrations.</p> <p><b>CO6-Creation:</b> Students will design and optimize conductometric experiments for complex acid-base and precipitation reactions, integrating digital tools for data analysis.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO3</b>  | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
| <b>CO6</b>                          | ✓   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |

# **Third Year B. Sc. Semester – V**

## **Minor Course (Practical)**

### **CHP-ME-501**

#### **Syllabus effective from June 2025**

1. Determine the normality of HCl by conductometric titration with strong base.
2. Determine the normality of CH<sub>3</sub>COOH by conductometric titration with strong base.
3. Conductometric titration of mixture of acids (HCl + CH<sub>3</sub>COOH) using strong base.
4. Conductometric titration of BaCl<sub>2</sub> vs K<sub>2</sub>CrO<sub>4</sub>.
5. Conductometric titration of AgNO<sub>3</sub> vs NaCl.

(Note: Perform minimum four practical)

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

## T. Y. B. Sc. – SEM – V CHEMISTRY (MINOR)

## PAPER - CH-ME-502

Applied Industrial Chemistry (Minor) (2 Credits Theory + 2 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-ME-502   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Applied Industrial Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to provide students with an understanding of the chemical processes involved in the fermentation industry, water pollution analysis, cosmetic chemistry, and small-scale chemical manufacturing.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn about fermentation processes for industrially important chemicals, water quality assessment techniques, the chemistry behind cosmetics and perfumes, and the formulation of everyday chemical products through small-scale preparations.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the fundamental concepts of fermentation, water pollution, cosmetics chemistry, and small-scale chemical preparations.</p> <p><b>CO2-Understanding:</b> Students will explain fermentation conditions, pollutant effects, chemical formulations in cosmetics, and principles behind small-scale manufacturing.</p> <p><b>CO3-Application:</b> Students will apply chemical knowledge to industrial processes, wastewater treatment, and formulation of commercial products like cosmetics and household items.</p> <p><b>CO4-Analysis:</b> Students will critically analyse industrial processes, assess environmental impact, and evaluate ingredient roles in cosmetic formulations.</p> <p><b>CO5-Evaluation:</b> Students will evaluate industrial methods, compare chemical processes, and assess product effectiveness and sustainability.</p> <p><b>CO6-Creation:</b> Students will design and develop eco-friendly products and optimize industrial processes to improve efficiency and sustainability.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO3</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           |             | ✓           |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |

# Third Year B. Sc. Sem. – V

## CH-ME-502 (Applied Industrial Chemistry)

Syllabus effective from June 2025

|                 |   |             |
|-----------------|---|-------------|
| <b>Unit – 1</b> |   |             |
|                 | <b>Fermentation Industry</b><br>Definition, condition favorable for fermentation process (pH, temperature, presence of other substances, absence of preservatives, concentration), manufacture of ethanol, citric acid, acetone and butanol, acetic acid, lactic acid from molasses, manufacture of penicillin-G  | <b>08 h</b> |
|                 | <b>Water Pollution</b><br>Types of water pollutants, Trace elements in water and their effects; Determination of BOD, COD, DO, Total hardness, Total dissolved solids, Ozone treatment process for wastewater.  | <b>07 h</b> |
| <b>Unit – 2</b> |   |             |
|                 | <b>Chemistry of cosmetics and perfumes</b><br>A general study including preparation and uses of the following: Hair dye, hair spray, shampoo, suntan lotions, face powder, lipsticks, talcum powder, nail enamel, creams (cold, vanishing and shaving creams), antiperspirants and artificial flavours. Essential oils and their importance in cosmetic industries with reference to Eugenol, Geraniol, sandalwood oil, eucalyptus, rose oil, 2-phenyl ethyl alcohol, Jasmone, Civetone, Muscone. | <b>08 h</b> |
|                 | <b>Some small scale preparation of</b><br>(1) Safety matches (2) Naphthalene balls (3) Wax candles (4) Shoe polish<br>(5) Writing/ fountain pen ink (6) Chalk crayons (7) Plaster of Paris  | <b>07 h</b> |

### Reference Books:

- 1) Shreve Chemical Process Industries 5 ed. George. T. Austin. Mag. Hill. Book Agency
- 2) Reigel's Industrial Chemistry Ed. By James A. Kent.
- 3) Unit Process in Organic Synthesis by D. H. Groggins.
- 4) The Chemical Process Industries by R. Norris Shreve; McGraw-Hill Book Company, Ltd.
- 5) An Introduction to Industrial Chemistry by Peter Wiseman, Applied Science Pub. Ltd. London.
- 6) Industrial Chemistry by Clerk Ranken; Andesite Press.
- 7) Industrial Chemistry by B. K. Sharma Goel Pub.
- 8) Quantitative Analysis by R.A.Day & A L Underwood, 6th ed. Pub. Prentice Hall of India ltd.
- 9) Vogel's Text Book Inorganic Quantitative Analysis, 6<sup>th</sup> ed.
- 10) Industrial safety management, by L.M. Desmukh, Tata Mc Graw Hill, New Delhi, 2006.
- 11) Industrial safety, Health & Environment management, Sunil S. Rao, R.K. Jain. Khanna Publishers, New Delhi, 2006.
- 12) E. Stocchi: Industrial Chemistry, Vol -I, Ellis Horwood Ltd. UK.
- 13) P.C. Jain, M. Jain: Engineering Chemistry, Dhanpat Rai & Sons, Delhi.
- 14) Sharma, B.K. & Gaur, H. Industrial Chemistry, Goel Publishing House, Meerut (1996).

**VEER NARMAD SOUTH GUJARAT UNIVERSITY****T. Y. B. Sc. – SEM – V CHEMISTRY Practical (MINOR)****PAPER - CHP-ME-502****Applied Industrial Chemistry (Minor) (2 Credits Theory + 2 Credits Practical)****As per NEP 2020****To be implemented from the Academic year 2025-26**

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-ME-502  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Applied Industrial Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to develop students' practical skills in organic estimation and chromatographic separation techniques, essential for quantitative and qualitative analysis in organic chemistry.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to determine the concentration and purity of organic compounds such as ketones, esters, and amino acids using titration and estimation techniques. Additionally, they will gain hands-on experience in separating amino acids using ascending paper chromatography, enhancing their analytical and laboratory skills.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental principles of organic estimation and chromatography, including separation techniques and quantitative analysis.</p> <p><b>CO2-Understanding:</b> Students will explain the principles behind organic estimation methods and the working of chromatographic techniques for amino acid separation.</p> <p><b>CO3-Application:</b> Students will apply various titrimetric and chromatographic techniques to determine the composition and purity of organic compounds.</p> <p><b>CO4-Analysis:</b> Students will analyse experimental data, interpret chromatographic results, and assess the accuracy and precision of organic estimations.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the reliability of different organic estimation techniques and compare chromatographic separation efficiencies.</p> <p><b>CO6-Creation:</b> Students will design optimized protocols for organic compound estimation and chromatography, considering accuracy and efficiency.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>  | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |

**Third Year B. Sc. Semester – V**  
**Minor Course (Practical)**  
**CHP-ME-502**  
**Syllabus effective from June 2025**

**ORGANIC ESTIMATION (Any Four)**

1. Determination of amount of Ketone (Acetone)
2. Determination of saponification value of an oil
3. Determination of percentage purity of Aspirin.
4. Determination of amount of Formaldehyde in given solution.
5. Determination of amount of Ethyl acetate in the given solution.
6. Determination of amount of Glycine in the given solution.

(Instead of Sample weighing, solution to be given)

**CHROMATOGRAPHY**

Chromatography separation of amino acid mixture by ascending paper chromatography. **(Any two)**

1. Glycine + Methionine
2. Alanine + Methionine
3. Alanine + Valine

[Subject Code-2503000505062001]

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**

**T. Y. B. Sc. – SEM – V CHEMISTRY (SEC)**

**PAPER - CH-SEC-501**

**Spectral and Industrial Chemistry (SEC) (1 Credits Theory + 1 Credits Practical)**

**As per NEP 2020**

**To be implemented from the Academic year 2025-26**

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-SEC-501   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Spectral and Industrial Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 01   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to provide students with fundamental knowledge of molecular spectroscopy and the chemistry of soaps and detergents, emphasizing their principles, applications, and environmental impact.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn the principles and applications of molecular spectroscopy, including NMR, ESR, and Raman spectroscopy, along with the chemistry, types, and cleansing mechanisms of soaps and detergents.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental principles of molecular spectroscopy, including different types of spectra and electromagnetic radiation.</p> <p><b>CO2-Understanding:</b> Students will explain the principles and applications of NMR, ESR, rotational, vibrational, and Raman spectroscopy, along with the mechanism of light scattering.</p> <p><b>CO3-Application:</b> Students will apply spectroscopic techniques to analyze molecular structures, calculate bond lengths, and interpret spectral data.</p> <p><b>CO4-Analysis:</b> Students will analyse the differences between various types of molecular spectra and evaluate the principles of single-beam spectrophotometers.</p> <p><b>CO5-Evaluation:</b> Students will compare the properties, environmental impacts, and cleansing actions of soaps and detergents based on their chemical composition.</p> <p><b>CO6-Creation:</b> Students will design and propose improved formulations for soaps and detergents, considering efficiency, environmental sustainability, and industrial applications.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO5</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |



[Subject Code-2503000505062002]

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

T. Y. B. Sc. – SEM – V CHEMISTRY Practical (SEC)

PAPER - CHP-SEC-501

Spectral and Industrial Chemistry (SEC) (1 Credits Theory + 1 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-SEC-501   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Spectral and Industrial Chemistry (Practical)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 01  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course aims to equip students with practical skills in spectrophotometric analysis, surfactant behavior studies, and soap formulation.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to analyse chemical compounds using spectrophotometry, study the effects of electrolytes on surfactants, determine critical micelle concentration, and assess soap quality through TFM measurement and liquid soap preparation.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>C01-Remembering:</b> Students will recall the principles of spectrophotometry and surfactant chemistry, including key concepts such as cloud point, critical micelle concentration (CMC), and total fatty matter (TFM).</p> <p><b>C02-Understanding:</b> Students will explain the interactions of electrolytes with surfactants, the chemistry of soap preparation, and the role of metal complexes in spectrophotometric analysis.</p> <p><b>C03-Application:</b> Students will apply spectrophotometric methods to determine <math>\text{NO}_2^-</math> and metal-ligand complexes, and analyse the impact of electrolytes on non-ionic surfactants.</p> <p><b>C04-Analysis:</b> Students will analyse experimental data to determine CMC values, assess soap quality based on TFM, and evaluate the cloud point behavior of surfactants.</p> <p><b>C05-Evaluation:</b> Students will compare the effectiveness of different spectrophotometric techniques, assess the influence of salts on surfactant properties, and evaluate soap formulations.</p> <p><b>C06-Creation:</b> Students will design and formulate liquid glycerine soap, optimizing composition for improved cleansing efficiency and environmental sustainability.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
| <b>C01</b>                          | ✓   | ✓           | ✓           |             |             | ✓           | ✓           | ✓           | ✓           |
| <b>C02</b>                          | ✓   |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
| <b>C03</b>                          | ✓   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
| <b>C04</b>                          | ✓   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
| <b>C05</b>                          | ✓   | ✓           | ✓           |             |             | ✓           | ✓           | ✓           | ✓           |
| <b>C06</b>                          | ✓   |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |

## Third Year B. Sc. Sem. – V

### CHP-SEC-501

#### Syllabus effective from June 2025

1. Spectrophotometric determination of  $\text{NO}_2^-$ .
2. Spectrophotometric determination of copper-ammonium complex.
3. Effect of electrolyte on cloud point of non-ionic surfactant.
4. CMC determination of ionic surfactant in the presence of salt.
5. Preparation of liquid glycerine soap.
6. Measurement of TFM (Total fatty matter) in soaps.

(Note: Perform any four experiments)

#### Reference Books:

- Handbook of Food Analysis: Physical Characterization and Nutrient Analysis by Leo M.L. Nollet, 3<sup>rd</sup> Edition, 2015, CRC Press.
- Food Adulteration and Safety by H.G. Ramachandran 1<sup>st</sup> Edition, 2016, Tata McGraw-Hill Education
- Advanced Practical Physical Chemistry by J. B. Yadav, 31<sup>st</sup> Edition, 2020, Goel Publishing House.
- Food Chemistry, by H.-D. Belitz, Werner Grosch, Peter Schieberle, 4<sup>th</sup> Revised and Extended Edition, 2009, Springer.
- Handbook of Food Adulteration and Safety, by B.K. Ghosh, 1<sup>st</sup> Edition, 2020, Random Publications.
- Modern Technology of Soaps, Detergents and Toiletries by P.K. Chattopadhyay, Revised Edition, 2013, NIIR Project Consultancy Services.
- Surfactants and Interfacial Phenomena by M. J. Rosen, and Joy T. Kunjappu, 4<sup>th</sup> Edition, 2012, Wiley-Interscience.
- Vogel's Quantitative Chemical Analysis by J. Mendham, R.C. Denney, J.D. Barnes, and M. Thomas, 6<sup>th</sup> Edition, 2000, Pearson Education.

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

## T. Y. B. Sc. – SEM – VI CHEMISTRY (MAJOR)

## PAPER - CH-MJ-601

## Inorganic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-601  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Inorganic Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course introduces molecular symmetry, inorganic reaction mechanisms, and organometallic compounds, emphasizing their structural and bonding aspects.  |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to classify molecules into point groups, understand ligand substitution mechanisms in metal complexes, and explore the bonding in key organometallic compounds.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the fundamental concepts of molecular symmetry, point groups, and organometallic compounds, including their classification and bonding.</p> <p><b>CO2-Understanding:</b> Students will explain symmetry elements, symmetry operations, ligand substitution mechanisms, and bonding in organometallic compounds.</p> <p><b>CO3-Application:</b> Students will classify molecules into point groups, apply symmetry rules to chemical structures, and determine reaction mechanisms in metal complexes.</p> <p><b>CO4-Analysis:</b> Students will analyse molecular structures using symmetry principles, construct multiplication tables, and evaluate the stability of organometallic compounds.</p> <p><b>CO5-Evaluation:</b> Students will compare different ligand substitution mechanisms, predict reaction pathways in inorganic complexes, and assess the bonding nature in organometallic compounds.</p> <p><b>CO6-Creation:</b> Students will construct symmetry multiplication tables, predict the behavior of organometallic compounds, and propose mechanisms for substitution reactions.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>   | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO5</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |

# CH-MJ-601 (Inorganic Chemistry)

Program Name: T. Y. B. Sc. Sem. – VI

Syllabus effective from June 2025

|                 |   |                                |
|-----------------|---|--------------------------------|
| <b>Unit – 1</b> |   |                                |
|                 | <b>Molecular Symmetry</b><br>Introduction and importance of symmetry, Symmetry elements and Symmetry operations, Classification of molecules in to point groups. Point group of simple molecules like CO <sub>2</sub> , HCl, H <sub>2</sub> O, NH <sub>3</sub> , BF <sub>3</sub> , [PtCl <sub>4</sub> ] <sup>-2</sup> , PF <sub>5</sub> , C <sub>6</sub> H <sub>6</sub> , C <sub>5</sub> H <sub>5</sub> <sup>-</sup> , Bromo benzene (C <sub>6</sub> H <sub>5</sub> Br), Cyclobutane, Boric acid (H <sub>3</sub> BO <sub>3</sub> ), Cis and Trans Dichoro ethylene (C <sub>2</sub> H <sub>2</sub> Cl <sub>2</sub> ), Staggered and Eclipsed Ethane (C <sub>2</sub> H <sub>6</sub> ). Law of multiplications, Construction of multiplication table for C <sub>2v</sub> , C <sub>3v</sub> , C <sub>2h</sub> . | <b>15 h</b>                    |
| <b>Unit – 2</b> |   |                                |
|                 | <b>[A] Metal Complexes (Inorganic Reaction Mechanism)</b><br>Reaction mechanisms of ligand substitution in octahedral complexes (i) SN <sup>1</sup> (ii) SN <sup>2</sup> , Acid hydrolysis & Base Hydrolysis-Redox (Single Electron Transfer) reactions, Substitution reaction without breaking M-L bond.<br><b>[B] Organo-metallic compounds</b><br>Definition, classification, Structure and bonding in ferrocene, dibenzene chromium, Zeise ion and gaseous dimethyl beryllium, Tetramethyl lead.  | <b>08 h</b><br><br><b>07 h</b> |

## Reference books:

- 1) Theoretical Inorganic chemistry by Day & Selbin, Affiliated East West
- 2) Advanced Inorganic Chemistry by Cotton and Wilkinson, John Wiley
- 3) Uni. Chemistry by B. H. Mahan
- 4) Structural Inorganic chemistry by A. F. Wells.
- 5) Chemical Bonding- an introduction By Rawal, Patel & Patel. Sugumar.
- 6) Basic Inorganic Chemistry by Cotton and Wilkinson
- 7) A Text book of Inorganic Chemistry by P.L. Soni
- 8) Introduction to Inorganic Chemistry by Durrant and Durrant
- 9) Modern Co-ordination Chemistry by R. Lewis and R.G. Wilkinson.
- 10) Inorganic Chemistry- Principles of structure and reactivity by J.E. Huhhey and E.A. Keiter.
- 11) Application of Group Theory to Chemistry by P. K. Bhattacharya., Himalaya Pub. House, Mumbai.

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – VI CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-601**

**Inorganic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

**As per NEP 2020**

**To be implemented from the Academic year 2025-26**

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-601  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Inorganic Chemistry (Practical)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course provides hands-on experience in gravimetric and volumetric analysis, focusing on the estimation of metal ions and alloys using classical analytical techniques.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will develop skills in gravimetric and volumetric methods for quantitative analysis, including metal ion precipitation, complexometric titrations, and purity determination of compounds.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the principles and procedures involved in gravimetric and volumetric estimation methods for metal ions and alloys.</p> <p><b>CO2-Understanding:</b> Students will explain the theoretical concepts of precipitation reactions, complexometric titrations, and the role of indicators in volumetric analysis.</p> <p><b>CO3-Application:</b> Students will perform gravimetric and volumetric estimations, apply stoichiometric principles, and analyse alloy compositions.</p> <p><b>CO4-Analysis:</b> Students will interpret experimental results, assess reaction completion, and determine sources of errors in gravimetric and volumetric methods.</p> <p><b>CO5-Evaluation:</b> Students will compare different titrimetric and gravimetric techniques, assess sample purity, and evaluate the accuracy and precision of their experimental findings.</p> <p><b>CO6-Creation:</b> Students will design experimental procedures for alloy and metal ion estimation, optimize reaction conditions, and propose modifications to improve accuracy.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |             |
|                                     | <b>CO3</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             | ✓           |
|                                     | <b>CO4</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |

# CHP-MJ-601 (Inorganic Chemistry)

Third Year B. Sc. Semester - VI

Syllabus effective from June 2025

## GRAVIMETRIC ESTIMATION OF (ANY TWO)

1.  $\text{Fe}^{+2}$  as  $\text{Fe}_2\text{O}_3$  from  $\text{Fe}(\text{NH}_4)_2\text{SO}_4 + \text{CuSO}_4$
2.  $\text{Ba}^{+2}$  as  $\text{BaSO}_4$  from  $\text{BaCl}_2 + \text{FeCl}_3$
3.  $\text{Al}^{+3}$  as  $\text{Al}_2\text{O}_3$  from  $\text{Al}_2(\text{SO}_4)_3 + \text{CuSO}_4$

## ESTIMATION OF ALLOY (ANY ONE)

1. **Brass-**  
Zinc as  $\text{Zn}_2\text{P}_2\text{O}_7$  (gravimetrically) & Copper by iodometrically (volumetric)
2. **German silver -**  
Nickel as  $\text{Ni}(\text{DMG})_2$  (gravimetrically) & Copper by iodometrically (volumetric)

## VOLUMETRIC EXERCISE (ANY FOUR)

1. To determine the percentage purity of potassium acid phthalate.
2. To determine the amount of Ammonium sulphate in the given solution
3. To determine the amount of Ferric by EDTA.
4. To determine the amount of Chromium by EDTA.
5. To determine the amount of Nickel with Magnesium by EDTA.
6. To determine the amount of Chloride by Mohr's method OR Absorption indicator.
7. To determine the amount of Bromide by Vohlard's method OR Absorption indicator.
8. To determine the percentage purity of  $\text{NaNO}_2$  /  $\text{KNO}_2$ .

[Subject Code-2003000300022001]

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**

**T. Y. B. Sc. – SEM – VI CHEMISTRY (MAJOR)**

**PAPER - CH-MJ-602**

**Organic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-602  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Organic Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course explores the intersection of Ayurveda and organic chemistry, emphasizing the chemical composition of Ayurvedic herbs and molecular transformations in organic compounds.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will understand the scientific basis of Ayurvedic principles, learn key molecular rearrangements, and study the structure and synthesis of important terpenoids.  |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the fundamental concepts of Ayurveda, organic chemistry, and the chemical composition of Ayurvedic herbs.</p> <p><b>CO2-Understanding:</b> Students will explain the scientific validation of Ayurvedic principles and their applications in organic chemistry.</p> <p><b>CO3-Application:</b> Students will analyse the mechanisms of molecular rearrangements, including Wagner-Meerwein, Pinacol-Pinacolone, Hoffmann, Curtius, and Beckmann rearrangements.</p> <p><b>CO4-Analysis:</b> Students will classify terpenoids, apply the isoprene rule, and determine structural aspects of <math>\alpha</math>-pinene and citral using analytical and synthetic evidence.</p> <p><b>CO5-Evaluation:</b> Students will compare molecular rearrangements, assess the significance of isoprenoids in natural product chemistry, and validate Ayurvedic chemical principles.</p> <p><b>CO6-Creation:</b> Students will propose future directions for integrating Ayurveda and organic chemistry, design synthesis strategies for terpenoids, and develop research-based applications.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
| <b>CO6</b>                          | ✓  | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |             |

# CH-MJ-602 (Organic Chemistry)

Program Name: T. Y. B. Sc. Sem. – VI

Syllabus effective from June 2025

|                 |  |             |
|-----------------|--|-------------|
| <b>Unit – 1</b> |  |             |
|                 | <b>Panchmahabhuta and organic compounds</b><br>Introduction to Ayurveda & Organic Chemistry.<br>Chemical composition of Ayurvedic herbs.<br>Scientific validation of Ayurvedic Principles.<br>Applications & Future directions.  | <b>15 h</b> |
| <b>Unit – 2</b> |  |             |
|                 | <b>(A) Molecular Rearrangements</b><br>Mechanism of rearrangements involving C to C migrations as illustrated by Wagner-Meerwein and Pinacol-Pinacolone rearrangements.<br>Mechanism of rearrangements involving C to N migrations as illustrated by Hoffmann, Curtius, and Beckmann rearrangements. | <b>8 h</b>  |
|                 | <b>(B) Terpenoids (Isoprenoids)</b><br>Their occurrence, classification, isoprene, and special isoprene rule, general methods to determine their structure, analytical and synthetic evidences for the structure of $\alpha$ -pinene & Citral.   | <b>7 h</b>  |

## Reference Books:

- Baldev Upadhyaya, Samskrta Sastrom ka Itihas, Chowkhambha, Varansi, 2010.
- D. M. Bose, S. N. Sen and B. V. Subbarayappa, Eds. A Concise History of Science of India, 2<sup>nd</sup> edition, University press, Hydrabad, 2010.
- Astangahrdaya Vol. 1 Sutrasthana and Sarirasthana, Translated by K. R. Srikantha Murthy, Vol 1 Krishnadas Academy, Varansi, 1991.
- Dharampal, Some aspects of earlier Indian Society and polity and their relevance today, New quest Publications, Pune, 1987.
- Dharampal, Indian science and technology in the eighteenth century; Some contemporary European accounts, Dharampal classic series, Rashtrottana Sahitya, Bengaluru, 2021
- Mechanism and Structure in organic chemistry-GouldeS.
- Reaction mechanism in organic chemistry by Mukhejee &Singh
- Principles of reaction mechanism in organic chemistry by Dharmaraha & Chawla
- Organic reaction mechanism by Bansal Tata Mac.Hill
- Organic Chemistry by Hendrickson, Cram &Hammond
- Organic Chemistry by Brown R.F.
- Organic Chemistry by Solomon W. Graham
- Principles of Organic Synthesis- R. O. C. Norman
- Basic Principles of Organic chemistry, by R. Y. Caserio, W. A.Benjamin
- May's Chemistry of synthetic Drugs by Dyson
- Synthetic Organic Chemistry by O. P. Agarwal
- Organic Chemistry by Morrison and Boyd
- Chemistry of organic Natural Product Vol. I & II by O. P.Agarwal
- Organic reactions & their mechanism by P. S. Kalsi, New age international publishers.

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – VI CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-602**

**Organic Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |  |             |             |             |             |             |             |             |             |
|-------------------------------------|--|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-602   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Organic Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02   |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours   |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026  |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course focuses on the systematic separation, identification, and purification of organic compounds to enhance students' practical skills in organic analysis.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn techniques for separating binary mixtures, identifying components, and performing crystallization and derivative preparation for purification.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall the principles of organic separation techniques, including crystallization and derivative preparation.</p> <p><b>CO2-Understanding:</b> Students will explain the methods used for separating binary mixtures containing acids, bases, phenols, and neutral compounds.</p> <p><b>CO3-Application:</b> Students will analyse and identify individual components of binary organic mixtures using suitable separation techniques.</p> <p><b>CO4-Analysis:</b> Students will classify organic compounds into different functional groups and determine their derivatives through purification techniques.</p> <p><b>CO5-Evaluation:</b> Students will compare separation techniques, assess the efficiency of different purification processes, and validate separation results.</p> <p><b>CO6-Creation:</b> Students will design and develop improved methods for organic separation, optimizing techniques for purification and characterization.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |  | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>   | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO2</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>   | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>   | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO6</b>   | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |

# CHP-MJ-602 (Organic Chemistry)

Third Year B. Sc. Semester - VI

Syllabus effective from June 2025

## ORGANIC SEPARATION

Separation of binary mixture, identification of the components and its crystallization & preparation of one derivative and its purification:

- ACID** : Benzoic acid, Salicylic acid, Phthalic acid, Cinnamic acid, Phenyl acetic acid
- BASE** : *o* - Nitroaniline, *m* - Nitroaniline, *p* - Nitroaniline, Aniline, *p* - Toluidine, *p* - Chloroaniline, Dimethylaniline, Diethylaniline, Diphenylamine  
(Not with Neutral)
- PHENOL** :  $\alpha$  - naphthol,  $\beta$  - naphthol, *o* - Nitro Phenol
- NEUTRAL** : **Aldehyde**: Benzaldehyde  
**Ketone**: Acetone, Methyl Ethyl ketone, Acetophenone  
**Ester**: Methyl acetate, Ethyl acetate,  
**Alcohol**: Methanol, Ethanol  
**Hydrocarbon**: Toluene, Anthracene, Naphthalene, Diphenyl  
**Nitro Hydrocarbon**: Nitrobenzene, *m* - Dinitro benzene  
**Halogenated hydrocarbon**: Chloroform, Carbon tetrachloride, Chlorobenzene, Bromobenzene, *p* - Dichlorobenzene  
**Amide**: Benzamide  
**Anilide**: Acetanilide

**NOTE:** Candidate should perform the analysis of at least 08 mixtures.

**VEER NARMAD SOUTH GUJARAT UNIVERSITY****T. Y. B. Sc. – SEM – VI CHEMISTRY (MAJOR)****PAPER - CH-MJ-603****Analytical Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-MJ-603   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Analytical Chemistry (Theory)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course introduces the principles and instrumentation of spectroscopic and chromatographic techniques essential for chemical analysis and identification.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn the fundamentals of Raman spectroscopy, UV-Vis spectrophotometry, gas chromatography, and high-performance liquid chromatography for qualitative and quantitative analysis.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental principles of spectroscopy, Raman spectroscopy, and chromatography techniques.</p> <p><b>CO2-Understanding:</b> Students will explain the theories, instrumentation, and applications of Raman spectroscopy, spectrophotometry, gas chromatography, and liquid chromatography.</p> <p><b>CO3-Application:</b> Students will apply spectrophotometric and chromatographic techniques for quantitative analysis of different ions such as <math>\text{Cu}^{2+}</math>, <math>\text{Fe}^{3+}</math>, and <math>\text{NO}_2^-</math>.</p> <p><b>CO4-Analysis:</b> Students will analyse chromatographic data to differentiate between packed and capillary columns, detector selection, and sample introduction techniques.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the efficiency of single-beam and double-beam spectrophotometers and compare qualitative and quantitative analysis methods in GC and HPLC.</p> <p><b>CO6-Creation:</b> Students will design experiments using spectrophotometry and chromatography to optimize analytical methods for chemical analysis.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |             |
|                                     | <b>CO3</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO4</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |             |

# CH-MJ-603 (Analytical Chemistry)

Program Name: T. Y. B. Sc. Sem. – VI

Syllabus effective from June 2025

| Unit – 1 |  |                                |
|----------|--|--------------------------------|
|          | <p><b>Spectroscopy</b><br/>Introduction of Raman spectroscopy (Indian Knowledge System), Basic principle, Theory of Raman spectroscopy, Instrumentation; Source, Sample handling (solid, liquid and gas samples).<br/>Different types of spectrum, Process involved in interaction with matter (Fluorescence, Phosphorescence), Components of Spectrophotometer-Sources, Grating and Prism as dispersing device, Sample handling, Detectors- Photo tube, Photomultiplier tube. Block diagram and working of single beam and double beam spectrophotometer. Terms involved in Beer's law (no derivation). Lambert-Beer's law and its derivation, Analysis of unknown by calibration curves method. Determination of <math>\text{Cu}^{+2}</math>, <math>\text{Fe}^{+3}</math>, <math>\text{NO}_2^{-1}</math>, using spectrophotometer. (Only principles), Problems based on quantitative analysis.</p> | <b>15 h</b>                    |
| Unit – 2 |  |                                |
|          | <p><b>Gas Chromatography</b><br/>Classification of chromatography, Principles of GC separation. Components of GC, Sample introduction system, Columns: Packed column Capillary Column (WCOT, SCOT), Carrier gas and its selection-stationary phases: Solid adsorbents, Inert supports (Selection criteria, Diatomaceous earths) and liquid stationary phases, Detectors: FID, TCD, Qualitative and quantitative analysis using GC.</p> <p><b>Liquid Chromatography</b><br/>Limitation of conventional liquid chromatography (no detail method), technique of HPLC, elementary idea about technique and layout diagrams of instrument, components of instrument of HPLC technique, Elementary idea of TLC.</p>  | <b>10 h</b><br><br><b>05 h</b> |

## Reference Books:

- Molecular diffraction of light by Sir C. V. Raman, Calcutta University Press, 1922.
- The Raman Effect: A unified treatment of the theory of Raman scattering by molecules by Derek A. Long
- Quantitative Analysis by R. A. Day & A. L. Underwood, 6<sup>th</sup> ed. Pub. Prentice Hall of India Ltd.
- Vogel's Text Book Inorganic Quantitative Analysis, 6th ed.
- Analytical Chemistry (Principles & Technique) by Lary G. Hargis
- Fundamental of Analytical Chemistry by Skoog D. A. & West D. M.
- Instrumental Methods of Analysis by B. K. Sharma
- Instrumental analysis by R. D. Braun Mc Graw Hill
- Analytical Chemistry by Gary Christian
- Modern Analytical Chemistry by David Harvey, McGraw Hill Higher Education
- College Analytical Chemistry, Mangaonkar, Teckchandani, Sathe, Ghalsasi, Jain, Himalaya Publishing House
- Analytical Chemistry by Alka L. Gupta, Pragati Prakashan
- Instrumental Methods of Chemical Analysis by H. Kaur, Pragati Prakashan

**VEER NARMAD SOUTH GUJARAT UNIVERSITY**  
**T. Y. B. Sc. – SEM – VI CHEMISTRY Practical (MAJOR)**

**PAPER - CHP-MJ-603**

**Analytical Chemistry (Major) (2 Credits Theory + 2 Credits Practical)**

**As per NEP 2020**

**To be implemented from the Academic year 2025-26**

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-MJ-603  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | Analytical Chemistry (Practical)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course provides hands-on experience in analysing reaction kinetics, surface tension, pH metry, conductometry, colorimetry, and refractometry to develop practical skills in chemical analysis.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will learn to determine reaction rates, dissociation constants, and concentrations of various compounds using instrumental techniques while understanding their applications in chemical research.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental concepts of reaction kinetics, surface tension, pH metry, conductometry, colorimetry, and refractometry.</p> <p><b>CO2-Understanding:</b> Students will explain reaction rate dependencies, dissociation constants, and the principles behind instrumental techniques like conductometry, pH metry, and colorimetry.</p> <p><b>CO3-Application:</b> Students will apply experimental techniques to determine reaction rates, dissociation constants, vanillin content, and the cleansing power of detergents.</p> <p><b>CO4-Analysis:</b> Students will analyse experimental data related to acid dissociation, Lambert-Beer's law, and refractometry for binary liquid mixtures.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the accuracy of experimental methods for determining reaction rates, weak acid constants, and specific refractivity.</p> <p><b>CO6-Creation:</b> Students will design and optimize experiments for titrations, conductometric analysis, and colorimetric verification of theoretical laws.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO3</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |             |

# CHP-MJ-603 (Analytical Chemistry)

Third Year B. Sc. Semester - VI

Syllabus effective from June 2025

1. To investigate the rate of reaction between  $\text{KBrO}_3$  and  $\text{KI}$ ,  $[a=b]$
2. To investigate the rate of reaction between  $\text{KBrO}_3$  and  $\text{KI}$ ,  $[a \neq b]$
3. **Surface tension:** To compare the cleansing power of two detergents by measuring surface tension of their solutions.
4. **pH metry:** To determine the dissociation constant of weak acid  $[\text{CH}_3\text{COOH}]$  by titration of weak acid and strong base  $[\text{NaOH}]$ .
5. **Conductometry:** To determine the amount of vaniline in the given sample of vanilla solution.
6. **pH metry:** To determine the amount of  $\text{Na}_2\text{CO}_3$  and  $\text{NaHCO}_3$  in the given mixture by using standard  $\text{HCl}$  solution.
7. **Conductometry:** To determine the amount of  $\text{CH}_3\text{COOH}$  in the given solution using  $\text{NH}_4\text{OH}$  solution.
8. **Colorimetry:** To determine the indicator constant of phenolphthalein.
9. **Colorimetry:** To verify Lambert-Beer's law for  $\text{KMnO}_4$  solution.
10. **Refractometry:** To determine the specific refractivity of the given liquids A, B and their mixtures containing 20%, 40%, 60% and unknown liquid by volume. [Benzene + Chloroform] OR [Benzene + Carbon tetra chloride]

**Note: Any seven practicals including TWO chemical kinetics should be performed by the students.**

## VEER NARMAD SOUTH GUJARAT UNIVERSITY

## T. Y. B. Sc. – SEM – VI CHEMISTRY (MINOR)

## PAPER - CH-ME-601

## General Chemistry (Minor) (2 Credits Theory + 2 Credits Practical)

As per NEP 2020

To be implemented from the Academic year 2025-26

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CH-ME-601   |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | General Chemistry (Theory)  |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 02 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course introduces students to detergents, environmental pollution, solution preparation, and industrial chemical processes, emphasizing their practical applications and impact.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will understand detergent classification, pollution control methods, solution preparation techniques, and key industrial processes with flow diagrams to develop applied chemical knowledge.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental concepts of detergents, pollution types, solution preparation, and industrial processes.</p> <p><b>CO2-Understanding:</b> Students will explain the principles of detergency, pollution control measures, and the theoretical basis of solution preparation and industrial synthesis.</p> <p><b>CO3-Application:</b> Students will apply their knowledge to classify detergents, analyse pollution effects, and prepare standard solutions.</p> <p><b>CO4-Analysis:</b> Students will analyse the interconversion of concentration units, evaluate industrial manufacturing processes, and study pollution control techniques.</p> <p><b>CO5-Evaluation:</b> Students will evaluate the environmental impact of pollutants, industrial chemical synthesis methods, and detergent properties.</p> <p><b>CO6-Creation:</b> Students will design flow diagrams for industrial processes and formulate solutions for pollution control and detergent formulation.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO3</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |             |

# Third Year B. Sc. Sem. – VI

CH-ME-601 (General Chemistry)

Syllabus effective from June 2025

| Unit – 1 |  |             |
|----------|--|-------------|
|          | <b>(A) Detergents</b><br>Introduction, Principles detergency, classification of surface-active agents, Anionic detergents, Cationic detergents, Nonionic detergents, Amphoteric detergents, Suds regulators, Builders and Additives.   | <b>08 h</b> |
|          | <b>(B) Environmental pollution</b><br>Introduction, types of Pollutions (1) Gaseous pollution in air, Acid rain, Green house Effect and ozone depletion (2) Radiation pollution cause, effect and control, (3) Noise Pollution and their effect and control (4) Oil pollution and their control.   | <b>07 h</b> |
| Unit – 2 |  |             |
|          | <b>(A) Preparation of solutions and standard solutions</b><br>Definitions of terms: Solute, Solvent, and Solution Composition of solution-normal solution, molar solution, molal solution, mole fraction, % solution, saturated, unsaturated and supersaturated solution and solubility. Effect of temp. on various units of concentration. Inter conversion of one unit into another unit. Preparation of solutions of some primary standard substances (e.g. Oxalic acid, succinic acid, KHP, $K_2Cr_2O_7$ , $As_2O_3$ ) | <b>10 h</b> |
|          | <b>(B) Industrial manufacturing process with flow diagram &amp; their uses</b> <ul style="list-style-type: none"><li>• Preparation of methanol from synthesis gas.</li><li>• Preparation of Isopropanol from propylene.</li><li>• Preparation of acetone from Isopropanol.</li><li>• Preparation of formaldehyde from methanol by oxidation dehydration process.</li><li>• Acetylene from natural gas.</li></ul>   | <b>05 h</b> |

## Reference Books:

- (1) Shreve Chemical Process Industries, 5ed., George. T. Austin. MacGraw Hill, Book Agency
- (2) Reigel's Industrial Chemistry, Ed. By James A. Kent
- (3) Unit Process in Organic Synthesis by D.H. Groggins
- (4) An Introduction to Industrial Chemistry, by Peter Wiseman, Applied Science Pub. Ltd. London.
- (5) Industrial Chemistry by B. K. Sharma, Goel Pub.
- (6) Quantitative Analysis by R. A. Day & A. L. Underwood, 6<sup>th</sup> ed. Pub. Prentice Hall of India ltd.
- (7) Vogel's Text Book Inorganic Quantitative Analysis, 6<sup>th</sup> ed.
- (8) Quantitative analysis by R.A. Day and A. L. Underwood
- (9) Elements of Analytical Chemistry by R. Gopalan; P.S. Subramanian and K. Rengarajan
- (10) Elementary Organic Spectroscopy by Y. L. Sharma
- (11) Organic Spectroscopy by B. K. Sharma
- (12) Environmental Chemistry by H. Kaur

**VEER NARMAD SOUTH GUJARAT UNIVERSITY****T. Y. B. Sc. – SEM – VI CHEMISTRY Practical (MINOR)****PAPER - CHP-ME-601****General Chemistry (Minor) (2 Credits Theory + 2 Credits Practical)****As per NEP 2020****To be implemented from the Academic year 2025-26**

|                                     |   |             |             |             |             |             |             |             |             |
|-------------------------------------|---|-------------|-------------|-------------|-------------|-------------|-------------|-------------|-------------|
| <b>Course Subject Code</b>          | CHP-ME-601  |             |             |             |             |             |             |             |             |
| <b>Subject Title</b>                | General Chemistry (Practical)   |             |             |             |             |             |             |             |             |
| <b>Credits</b>                      | 02  |             |             |             |             |             |             |             |             |
| <b>Teaching per week</b>            | 04 Hours  |             |             |             |             |             |             |             |             |
| <b>Effective from</b>               | 2025-2026   |             |             |             |             |             |             |             |             |
| <b>Purpose of Course</b>            | This course provides hands-on experience in titrimetric and instrumental analytical techniques for quantitative chemical analysis, essential for accurate determination of chemical substances.   |             |             |             |             |             |             |             |             |
| <b>Objective of Course</b>          | Students will develop skills in titrimetric methods and instrumental techniques like pH metry, conductometry, colorimetry, and surface tension measurement to analyse chemical compounds effectively.   |             |             |             |             |             |             |             |             |
| <b>Course Outcomes</b>              | <p><b>CO1-Remembering:</b> Students will recall fundamental principles of titrimetric analysis, EDTA titrations, Mohr's method, and instrumental techniques.</p> <p><b>CO2-Understanding:</b> Students will explain the underlying principles behind titrimetric methods and instrumental techniques like pH metry, conductometry, colorimetry, and surface tension measurement.</p> <p><b>CO3-Application:</b> Students will apply their knowledge to determine the purity of various compounds and analyse solutions using different analytical techniques.</p> <p><b>CO4-Analysis:</b> Students will analyse data obtained from titrations and instrumental methods to evaluate unknown concentrations and validate experimental results.</p> <p><b>CO5-Evaluation:</b> Students will evaluate different analytical techniques based on accuracy, precision, and applicability for various chemical analyses.</p> <p><b>CO6-Creation:</b> Students will develop experimental strategies for titrimetric and instrumental methods to solve analytical chemistry problems.</p> |             |             |             |             |             |             |             |             |
| <b>Mapping Between COs and PSOs</b> |   | <b>PSO1</b> | <b>PSO2</b> | <b>PSO3</b> | <b>PSO4</b> | <b>PSO5</b> | <b>PSO6</b> | <b>PSO7</b> | <b>PSO8</b> |
|                                     | <b>CO1</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO2</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |             |
|                                     | <b>CO3</b>  | ✓           |             | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |
|                                     | <b>CO4</b>  | ✓           | ✓           | ✓           | ✓           |             | ✓           | ✓           |             |
|                                     | <b>CO5</b>  | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           | ✓           |             |
|                                     | <b>CO6</b>  | ✓           | ✓           | ✓           |             | ✓           | ✓           | ✓           |             |

## Third Year B. Sc. Semester - VI

### CHP-ME-601

Syllabus effective from June 2025

#### TITRIMETRIC METHODS (Any Four)

1. To determination the amount of Bismuth by EDTA
2. To determination the amount of Nickel by EDTA
3. To determination the amount of Chloride by Mohr's method.
4. To determine the percentage purity of  $\text{NaNO}_2/\text{KNO}_2$
5. To determine the percentage purity of potassium acid phthalate.
6. To determine the percentage purity of Glycine.
7. To determine the percentage purity of Vitamin - C.

#### INSTRUMENTAL EXERCISE (Any three)

1. **pH Metry:** To determine the amount of HCl in given by Std. NaOH.
2. **Conductometry:** To determine the amount of HCl in given by Std. NaOH.
3. **Colourimetry:** To verify Lambert – Beer's law for  $\text{KMnO}_4$  solution.
4. **Surface Tension:** To compare the cleansing power of two detergents by measuring surface tension of their solutions.